

1. Give a brief over-view of each scientists claim-to-fame regarding our knowledge of the atom:
 - a. Aristotle
 - b. Democritus
 - c. Dalton
 - d. Thomson
 - e. Rutherford
 - f. Bohr

2. Name the three subatomic particles; give their charge; tell their location in the atom

3. Fill out the subatomic chart below:

Isotope Name	Atomic #	#P	#N	#E	Atomic Mass	Isotope symbol
V-50						
	24		28			
				19	39	
		37	48			

4. Calculate the weighted atomic mass of the element below:

The element has 2 isotopes: X and Y. Isotope X has a mass of 106.905 and is found 42% abundance in nature. Isotope Y has a mass of 108.905 and is found 58% abundance. What is the weighted atomic mass? What element is this?